Multiple Choice

1. The process of covering Iron with a protective layer of Zinc is referred to as
   1. Alloy
   2. Galvanization
   3. Sacrificial Anode
   4. Cathodic Protection
2. The process of attaching a more reactive metal to an iron object that will act as an Anode and prevent the corrosion of the Iron object is referred to as
   1. Galvanization
   2. Sacrificial Anode
   3. Cathodic Protection
   4. None of the above
3. Referring to the corrosion of Iron, which of the following is the Anode
   1. Iron
   2. Oxygen and Water
   3. Both a and b
   4. None of the Above

Short Answer

1. Using your own words, explain why stainless steel is a good method of preventing Iron corrosion.
2. You just bought a used patio set from your neighbour in great condition with no rust marks in the summer. Explain why after the winter, rust appeared on the set but it did not appear during the summer.

ANSWERS:

Multiple Choice:

1. B
2. C
3. A

Short Answer:

1. Stainless steel is an alloy of multiple metals. This allows for stainless steel to be much more resistant to corrosion than pure iron which is why it makes for a good method of preventing Iron corrosion.
2. Rust appeared after the winter because the metal was exposed to the Oxygen in the air and the water from the snow for a very long period of time. Rust did not appear in the summer time because the metal was exposed to a cool and dry environment.